Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

ACTIVITY SERIES OF METALS.

1. Explain why the following properties….. are suitable for gold jewellery making.
2. Explain why calcium is not found alone as an element in nature.
3. (2 marks)

3. Why can zinc be used as a coating to stop iron from rusting? (2 marks)

4. Explain why tin cans will rust when scratched but galvanised iron will not.

(2 marks)

5. Why does sodium have to be stored under paraffin oil? (1 mark)

6. When copper is used as a decorative item it is covered with a lacquer after being

polished. What is the reason for this? (1 mark)

7. Sacrificial anodes are used on aluminium boats.

a. Explain what is meant by a sacrificial anode. (1 mark)

b. Describe the purpose of the anode and how it works. (1 mark)

8. If copper metal is placed in a solution containing silver, what will happen? (1 mark)

**References = 2 marks. Therefore they must be correct.**

What is paraffin oil (Uses, history etc.)? It is often used to cover sodium. Please explain why sodium is stored in paraffin oil.

Discuss reactivity/oxidation.

6. What colour is copper when its unexposed, what colour dopes copper turn when it is exposed. Considering the answers to A/B why might you put a protective coating of lacquer on this solid.

7. Pictured are sacrificial anodes. Why would they be placed in this location in relation to the other metal described in the picture below.

What is a sacrificial anode?

7. What is a sacrificial anode? What is the purpose

(hint: